

Name _____ Date _____ Period _____



Chemical Reactions Ws #6: Classifying Reactions

Directions: Write a balanced equation and indicate the reaction type for each equations.

SR = single replacement DR = double replacement
C = Combination D = Decomposition

- _____ 1. Aluminum nitrate solution reacts with sodium hydroxide solution to yield sodium nitrate solution and a precipitate of aluminum hydroxide
- _____ 2. Solid potassium chlorate is heated to yield solid potassium chloride and oxygen gas.
- _____ 3. A solution of phosphoric acid reacts with magnesium hydroxide solution to give a precipitate of magnesium phosphate and liquid water.
- _____ 4. In the presence of heat, ammonium nitrite solid will yield nitrogen gas and liquid water.
- _____ 5. Ammonia gas reacts with oxygen gas to give nitrogen monoxide gas and water gas.
- _____ 6. Aqueous barium chloride reacts with aqueous sodium sulfate to give sodium chloride solution and a precipitate of barium sulfate.
- _____ 7. A solution of iron III bromide reacts with aqueous ammonium sulfide to give solid iron III sulfide and aqueous ammonium bromide.
- _____ 8. Solid calcium oxide combines with solid diphosphorus pentoxide to produce solid calcium phosphate
- _____ 9. An aqueous solution of magnesium chloride combines with an aqueous solution of silver nitrate to produce aqueous magnesium nitrate and solid silver chloride.
- _____ 10. Solid aluminum hydroxide and acetic acid solution combine to produce aqueous aluminum acetate and liquid water.
- _____ 11. Solid iron reacts with solid silver acetate to yield iron II acetate solution plus beautiful solid silver.
- _____ 12. Calcium hydroxide solution reacts with phosphoric acid to yield the salt precipitate calcium phosphate and water.
- _____ 13. Sodium hydroxide solution reacts with sulfuric acid to give the soluble salt sodium sulfate and water
- _____ 14. Magnesium nitrate (aq) + sulfuric acid (aq) → magnesium sulfate (aq) + nitric acid (aq)
- _____ 15. Cadmium III phosphate (aq) + Ammonium sulfide (aq) → cadmium III sulfide (aq) + ammonium phosphate (aq)