Carboxylic Acid Questions Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ # \_\_\_\_\_

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| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Top of Form   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | The correct structure for potassium benzoate would be:   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img442.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img443.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img444.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img445.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **2.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | The pH of an aqueous solution of acetic acid (vinegar) would be:   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | 7 |  |  |  |  | | --- | --- | --- | |  | b. | less than 7 |  |  |  |  | | --- | --- | --- | |  | c. | greater than 7 |  |  |  |  | | --- | --- | --- | |  | d. | 14 | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **3.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Which of the following is considered to be a fatty acid?   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img446.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img447.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img448.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img449.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **4.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Pictured below is methanoic acid. What is the common name for this carboxylic acid? /ilrn/books/beinl8/images/img450.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | methic acid |  |  |  |  | | --- | --- | --- | |  | b. | acetic acid |  |  |  |  | | --- | --- | --- | |  | c. | formic acid |  |  |  |  | | --- | --- | --- | |  | d. | carbonic acid | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **5.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | What is the IUPAC name for the following molecule? /ilrn/books/beinl8/images/img451.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | *o*-chlorobenzoic acid |  |  |  |  | | --- | --- | --- | |  | b. | *m*-chlorobenzoic acid |  |  |  |  | | --- | --- | --- | |  | c. | *p*-chlorobenzoic acid |  |  |  |  | | --- | --- | --- | |  | d. | carboxychlorobenzene | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **6.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Predict the product from the reaction shown below. /ilrn/books/beinl8/images/img452.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img453.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img454.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img455.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img456.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **7.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Which of the following carboxylic acids is the most soluble in water?   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img457.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img458.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img459.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img460.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **8.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Which of the carboxylic acids listed would be used to make polyethylene terephthalate (PET)? /ilrn/books/beinl8/images/img461.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img462.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img463.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img464.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img465.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **9.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Which of the following is the most soluble in water?   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img466.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img467.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img468.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img469.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **10.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Which of the molecules shown below is a carboxylic acid?   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img470.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img471.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img472.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img473.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **11.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Shown below is acetic acid. This would be classified as: /ilrn/books/beinl8/images/img474.pngAcetic Acid   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | a monoprotic acid |  |  |  |  | | --- | --- | --- | |  | b. | a diprotic acid |  |  |  |  | | --- | --- | --- | |  | c. | a triprotic acid |  |  |  |  | | --- | --- | --- | |  | d. | a tetraprotic acid | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **12.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | In the ionization of acetic acid in water, the conjugate base would be: /ilrn/books/beinl8/images/img475.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | CH3CO2H |  |  |  |  | | --- | --- | --- | |  | b. | H2O |  |  |  |  | | --- | --- | --- | |  | c. | CH3CO2- |  |  |  |  | | --- | --- | --- | |  | d. | H3O+ | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **13.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Carboxylic acids can be prepared by the oxidation of which class of alcohols?   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | quaternary |  |  |  |  | | --- | --- | --- | |  | b. | tertiary |  |  |  |  | | --- | --- | --- | |  | c. | secondary |  |  |  |  | | --- | --- | --- | |  | d. | primary | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **14.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | The IUPAC name for the compound below would be: /ilrn/books/beinl8/images/img476.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | 4-bromo-2-chlorobenzoic acid |  |  |  |  | | --- | --- | --- | |  | b. | *o-*chloro-*p-*bromobenzoic acid |  |  |  |  | | --- | --- | --- | |  | c. | 1-bromo-3-chlorobenzoic acid |  |  |  |  | | --- | --- | --- | |  | d. | 1-chloro-5-bromobenzoic acid | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **15.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Which of the molecules below is a structural isomer of propanoic acid? /ilrn/books/beinl8/images/img477.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img478.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img479.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img480.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img481.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **16.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | The final product of the reaction below would be: /ilrn/books/beinl8/images/img482.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img483.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img484.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img485.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img486.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **17.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Which of the following carboxylic acids would be least soluble in water?   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img487.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img488.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img489.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img490.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **18.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | The reaction between acetic acid and sodium hydroxide would produce: /ilrn/books/beinl8/images/img491.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | NaCl and H2O |  |  |  |  | | --- | --- | --- | |  | b. | Nothing, acetic acid and sodium hydroxide do not react. |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img492.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img493.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **19.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | The correct structure for 2-hydroxypropanoic acid would be:   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | /ilrn/books/beinl8/images/img494.png |  |  |  |  | | --- | --- | --- | |  | b. | /ilrn/books/beinl8/images/img495.png |  |  |  |  | | --- | --- | --- | |  | c. | /ilrn/books/beinl8/images/img496.png |  |  |  |  | | --- | --- | --- | |  | d. | /ilrn/books/beinl8/images/img497.png | |  | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  | | --- | | **20.** | | |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | The IUPAC name for the molecule shown below would be: /ilrn/books/beinl8/images/img498.png   |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | |  |  |  | | --- | --- | --- | |  | a. | sodium benzoic acid |  |  |  |  | | --- | --- | --- | |  | b. | sodium benzoate |  |  |  |  | | --- | --- | --- | |  | c. | potassium benzoate |  |  |  |  | | --- | --- | --- | |  | d. | benzoic sodium | |  | | |   Bottom of Form |

http://webquiz.ilrn.com/media/img/common/onepixel.gif