Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Class \_\_\_\_\_\_\_\_\_\_\_\_\_\_

***The Mole and Volume***

For gases of STP (273K and I atm pressure), one mole occupies a volume of 22.4 L. What volume will the following quantities of gases occupy of STP?

1. 1.00 mole of H2
2. 3.20 moles of O2
3. 0.750 mole of N2
4. 1.75 moles of CO2
5. 0.50 mole of NH3
6. 5.0 g of H2
7. 100 g of O2
8. 28.0 g of N2
9. 60 g of CO2
10. 10 g of NH3